

8. *Statistical Mechanics and Thermodynamics* (Spring 2007)

Consider a large number of photons, in thermal equilibrium at temperature T , inside a volume V . The photons are in vacuum and can have two possible directions of polarization.

(a) Let $u(\omega, T) d\omega$ be the mean energy per unit volume of photons in the frequency range between ω and $\omega + d\omega$. Find $u(\omega, T)$.

(b) What is the temperature dependence of the total energy density u_0 ?

Note: The constants of proportionality will involve a definite integral which you need not solve explicitly – find temperature dependence only.

$$a) \quad u_0 = \frac{1}{V} \langle U \rangle = \frac{1}{V} \int_0^{\infty} \epsilon f(\epsilon) \rho(\epsilon) d\epsilon \quad \epsilon = pc = \hbar\omega, \quad f(\epsilon) = \frac{1}{e^{\beta\epsilon} - 1}$$

$$\rho = \frac{d\Omega}{d\epsilon} = 2 \frac{1}{h^3} \frac{dV_{\text{phase}}}{d\epsilon} \quad (\text{where the 2 comes from polarization multiplicity})$$

$$dV_{\text{phase}} = (\int d^3r) (\int p^2 d\Omega_p) dp = V 4\pi p^2 dp = V 4\pi \left(\frac{\hbar}{c}\right)^3 \omega^2 d\omega$$

since $p = \frac{\hbar}{c}\omega$

$$\Rightarrow \rho = 2 \cdot 4\pi V \left(\frac{1}{2\pi c}\right)^3 \omega^2 \frac{d\omega}{d\epsilon} = \frac{V}{\pi^2 c^3} \omega^2 \frac{d\omega}{d\epsilon}$$

$$\Rightarrow u_0 = \frac{1}{V} \int_0^{\infty} \hbar\omega \left(\frac{1}{e^{\beta\epsilon} - 1}\right) \frac{V}{\pi^2 c^3} \omega^2 \frac{d\omega}{d\epsilon} d\epsilon$$

$$= \frac{\hbar}{\pi^2 c^3} \int_0^{\infty} \frac{\omega^3}{e^{\beta\hbar\omega} - 1} d\omega = \int_0^{\infty} u(\omega, T) d\omega$$

$$\Rightarrow u(\omega, T) = \frac{\hbar}{\pi^2 c^3} \frac{\omega^3}{e^{\hbar\omega/kT} - 1} \quad (\text{Planck's law})$$

$$b) \quad u_0 \sim \int_0^{\infty} \frac{\omega^3}{e^{\beta\hbar\omega} - 1} d\omega \sim \left(\frac{1}{\beta}\right)^4 \int_0^{\infty} \frac{x^3 dx}{e^x - 1}$$

where $x = \beta\hbar\omega$
so $\omega \sim \frac{1}{\beta} x$

$$\sim T^4 \quad (\sim \text{Stefan-Boltzmann Law})$$